LIST OF TABLES

CHAPTERS		TABLE CAPTIONS	PAGE NO.
CHAPTER II	TableII.a	Properties of Cyclodextrins	29
CHAPTER IV	**Table1**	Density (ρ), viscosity (η) and relative permittivity (ε) of solvent mixtures at 298.15 K[a]	99
	Table2	The concentration (c) and molar conductance (Λ) of $LiAsF_6$ in different mass fraction of EG (w_1) in MeOH at 298.15 K[a]	100
	Table3	Limiting molar conductance (Λo), association constant (K_a), co-sphere diameter (R) and standard deviations of experimental Λ (δ) of $LiAsF_6$ in different mass fraction of EG (w_1) in MeOH at 298.15 K[a]	102
	Table4	Walden product ($\Lambda_o.\eta$) and Gibb's energy change (ΔG^0) of $LiAsF_6$ in different mass fraction of ethylene glycol (w_1) in methanol at 298.15 K[a]	103
	Table5	Ionic limiting molar conductance ($\lambda_0^{\pm}$), ionic Walden product ($\lambda_0^{\pm}\eta$), crystallographic radii (r_c) and Stoke's radii (r_s) of Li+ and AsF_6- ion in different mass fraction of EG (w_1) in MeOH at 298.15 K[a]	104
	Table6	Experimental values of densities (ρ), apparent molar volume (φ_V), viscosites (η), of LiAsF6 in different mass fraction of EG (w_1) in MeOH at 298.15 K[a]	107
	Table7	Limiting apparent molar volume (ϕ_V^0), experimental slopes (S_V^*), *A*-, *B*- coefficients of $LiAsF_6$ in different mass fraction of EG (w_1) in MeOH at 298.15 K	108
CHAPTER V	**Table1**	Change in chemical shifts (ppm) of the H3 and H5 protons of cyclodextrin molecules in four different host-guest complexes in D_2O at 298.15 K[a].	118
	Table2	Values of surface tension (γ) and at the break point with corresponding concentration of aqueous α and β-cyclodextrins at 298.15 K[a]	121

List of Tables

CHAPTERS	TABLE CAPTIONS		PAGE NO.
	Table3	Association constants (K_a), Gibb's free energy, enthalpy and entropy of various ionic liquid-cyclodextrin systems	123
	Table4	The observed peaks at different m/z with corresponding ions for the solid inclusion complexes	126
	Table5	Limiting apparent molar volume (φv^0), Experimental slope (Sv*), Viscosity B and A coefficient of [BMIm]Cl and [HMIm]Cl in different concentration of aqueous α-CD and β- CD solvent mixtures at 298.15K[a]	129
	TableS1	1H-NMR spectra of [BMIm]Cl, [HMIm]Cl, α-CD, β-CD, and [BMIm]Cl+α-CD, [BMIm]Cl+β-CD, [HMIm]Cl+α-CD, [HMIm]Cl+β-CD complexes	132
	TableS2	Surface Tension (γ) values of aqueous [BMIm]Cl and [HMIm]Cl with α-CD and β-CD at 298.15 K[α]	133
	TableS3	Data for the conductivity study of aqueous [BMIm]Cl+β-CD and [BMIm]Cl+β-CD system (concentration of stock solution of [BMIm]Cl and [HMIm]Cl = 10mM, concentration of stock solution of β-CD = 10mM) at 293.15K[a], 303.15K[a], 313.15K[a]	134
	TableS4	Data for the conductivity study of aqueous [BMIm]Cl+α-CD and [HMIm]Cl+α-CD system (concentration of stock solution of [BMIm]Cl and [HMIm]Cl = 10mM, concentration of stock solution of α-CD = 10mM) at 293.15K[a], 303.15K[a], 313.15K[a]	135
	TableS5	Values of conductivity (k) at the break point with corresponding concentration of aqueous α and β-cyclodextrins at 298.15 K[a], 298.15 K[a], 298.15 K[a]	135
	TableS6	Data obtained from FT-IR spectroscopic study of α-CD, β-CD, [BMIm]Cl, [HMIm]Cl, α-CD+[BMIm]Cl, β-CD+[BMIm]Cl, α-CD+[HMIm]Cl, β-CD+[HMIm]Cl	136
	TableS7	Experimental values of density (ρ) and viscosity (η) in deferent mass fractions of aq. α-CD and β-CD	137

List of Tables

CHAPTERS	TABLE CAPTIONS		PAGE NO.
	TableS8	Experimental values of densities (ρ) and viscosities (η) with varying concentration in different mass fractions of aq. α-CD and β-CD [BMIm]Cl at 298.15K[a]	137
	TableS9	Experimental values of densities (ρ) and viscosities (η) with varying concentration in different mass fractions of aq. α-CD and β-CD for [HMIm]Cl at 298.15 K[a]	138
CHAPTER VI	**Table1**	Values of surface tension (γ) and at the break point with corresponding concentration of α –CD , β-CD for [BMPy]PF_6 at 298.15 K[a]	154
	Table2	Association constants obtained by the Benesi–Hildebrand method (K_a)from UV-Vis study and corresponding thermodynamic parameters and stoichiometry of [BMPy]PF_6& α-CD and [BMPy]PF_6& β-CD inclusion complexes at 293.15K[a], 303.15K[a] at 313.15K[a].	156
	Table3	Association constants obtained by the nonlinear program (K_a^c) from conductance study and corresponding thermodynamic parameters and stoichiometry of [BMPy]PF_6& α-CD and [BMPy]PF_6& β-CD inclusion complexes at 293.15K[a], 303.15K[a] and 313.15K[a].	157
	TableS1	Data for Job's Plot performed by UV-Vis spectroscopy for [BMPy]PF_6-β-CD system at 298.15 K[a]	163
	TableS2	Data for Job's Plot performed by UV-Vis spectroscopy for [BMPy]PF_6-α-CD system at 298.15 K[a]	163
	TableS3	1H-NMR spectra of [BMPy]PF_6, α-CD, β-CD, and [BMPy]PF_6+α-CD, [BMPy]PF_6+β-CD complexes	164
	TableS4	Change in chemical shifts (ppm) of the H3 and H5 protons of cyclodextrin molecule in two different host-guest complexes in D_2O at 298.15 K[a]	164
	TableS5	Surface Tension (γ) values of α-CD& [BMPy]PF_6, β-CD & [BMPy]PF_6 Complex at 298.15 K[a]	165

List of Tables

CHAPTERS	TABLE CAPTIONS		PAGE NO.
	TableS6	Data for the conductivity study of aqueous [BMPy]PF_6+α-CD and [BMPy]PF_6+β-CD system (concentration of stock solution of IL = 10mM, concentration of stock solution of CD = 10mM at 293.15K[a], 303.15K[a], 313.15K[a]	166
	TableS7	Values of conductivity (κ) at the break point with corresponding concentrations of [BMPy]PF_6 and CDs at 298.15 K[a]	167
	TableS8	Data for the Benesi-Hildebrand double reciprocal plot performed by UV-VIS spectroscopic study for [BMPy]PF_6-α-CD and [BMPy]PF_6-β-CD systems at 293.15 K[a]	167
	TableS9	Data for the Benesi-Hildebrand double reciprocal plot performed by UV-VIS spectroscopic study for [BMPy]PF_6-α-CD and [BMPy]PF_6-β-CD systems at 303.15 K[a]	168
	TableS10	Data for the Benesi-Hildebrand double reciprocal plot performed by UV-VIS spectroscopic study for [BMPy]PF_6-α-CD and [BMPy]PF_6-β-CD systems at 303.15K[a]	169
	TableS11	Data obtained from FT-IR spectroscopic study of α-CD, β-CD, [BMPy]PF_6, α-CD +[BMPy]PF_6, β-CD+[BMPy]PF_6.	170
	TableS12	The observed peaks at different m/z with corresponding ions for the solid inclusion complexes.	171
CHAPTER VII	**Table1**	Change in chemical shifts (ppm) of the H3 and H5 protons of cyclodextrin molecules in two different host-guest complexes in D_2O at 298.15 K[a]	180
	Table2	Values of surface tension (γ) at the break point with corresponding concentrations of TBPMS and CD at 298.15 K[a]	184
	Table3	Values of conductivity (κ) at the break point with corresponding concentrations of TBPMS and CD at 298.15 K[a]	185
	Table4	The observed peaks at different m/z with corresponding ions for the solid inclusion complexes	188

List of Tables

CHAPTERS		TABLE CAPTIONS	PAGE NO.
	Table5	Association constants (K_a), Gibb's free energy, enthalpy and entropy of various ionic liquid-cyclodextrin systems	189
	TableS1	1H-NMR spectra of TBPMS, α-CD, β-CD and TBPMS+α-CD, TBPMS+β-CD complexes	192
	TableS2	Surface Tension (γ) values of aqueous TBPMS with α-CD and β-CD at 298.15 K[a]	192
	TableS3	Data for the conductivity study of aqueous TBPMS+α-CD and TBPMS+β-CD system (concentration of stock solution of TBPMS = 10mM, concentration of stock solution of CD = 10mM at 293.15K[a], 303.15K[a], 313.15K[a]	193
	TableS4	Data obtained from FT-IR spectroscopic study of α-CD, β-CD, TBPMS and the complexes of α-CD & TBPMS, β-CD & TBPMS	194
CHAPTER VIII	**Table1**	Values of surface tension (γ) and at the break point with corresponding concentration of aqueous α and β-cyclodextrins at 298.15 K[a]	207
	Table2	Limiting apparent molar volume (φ_v^0), Experimental slope (S_v^*), Viscosity B and A coefficient of PTX and 2-PAM in different concentration of aqueous α-CD and β- CD solvent mixtures.	215
	TableS1	Surface Tension (γ) values of aqueous PTX and 2-PAM with α-CD and β-CD at 298.15 K[a]	219
	TableS2	Data for Job's Plot performed by UV-Vis spectroscopy for aqueous Pentoxifylline-α-CD system at 298.15 K[a]	219
	TableS3	Data for Job's Plot performed by UV-Vis spectroscopy for aqueous Pentoxifylline-β-CD system at 298.15 K[a]	220
	TableS4	Data for Job's Plot performed by UV-Vis spectroscopy for aqueous Pentoxifylline-β-CD system at 298.15 K[a]	220
	TableS5	Data for Job's Plot performed by UV-Vis spectroscopy for aqueous Pralidoxim-β-CD system at 298.15 K[a]	221
	TableS6	Data for the Benesi-Hildebrand double reciprocal plot performed by UV-VIS spectroscopic study for aqueous Drug-CD systems at 298.15 K[a]	221

List of Tables

CHAPTERS		TABLE CAPTIONS	PAGE NO.
	TableS7	1H-NMR spectra of PTX, 2-PAM, α-CD, β-CD, and PTX+α-CD,PTX+β-CD, 2-PAM+α-CD, 2-PAM+β-CD complexes	222
	TableS8	Change in chemical shifts (ppm) of the H3 and H5 protons of cyclodextrin molecules in four different host-guest complexes in D_2O at 298.15 K[a]	223
	TableS9	Data obtained from FT-IR spectroscopic study of α-CD, β-CD, PTX, 2-PAM, α-CD+PTX, β-CD+PTX, α-CD+2-PAM, β-CD+2-PAM	223
	TableS10	Experimental values of density (ρ) and viscosity (η) in deferent mass fractions of aq. α-CD and β-CD	224
	TableS11	Experimental values of densities (ρ) and viscosities (η) with varying concentration in different mass fractions of aq. α-CD with drug molecules.	224
CHAPTER IX	**Table1**	The values of density (ρ_0), viscosity (η_0) and refractive index (n_D) of Li-Ox, Na-Ox, K-Ox at different mass fractions of aqueous UA solution at 293.15K[a], 303.15 K[a], 313.15 K[a]	233
	Table2	The values of density (ρ), viscosity (η) of Li-Ox, Na-Ox, K-Ox at different mass fractions of aqueous UA solution at 293.15K[a], 303.15 K[a], 313.15 K[a]	234
	Table3	The values of apparent molar volume (φ_V), limiting apparent molar volume (φ_V^0), S_V^* of Li-Ox, Na-Ox, K-Ox at different mass fractions of aqueous UA solution at 293.15K[a], 303.15 K[a], 313.15 K[a]	236
	Table4	The values of a_0, a_1, a_2, φ_E^0, $(\delta\varphi_E^0/\delta T)_P$ Li-Ox, Na-Ox, K-Ox at different mass fractions of aqueous UA solution at 293.15K to 313.15K	237
	Table5	Values of viscosity A and B-coefficient for Li-Oxalate, Na-Oxalate, K-Oxalate against different mass fraction of aqueous UA (w_1) solution	240
	Table6	The values of refractive index (n_D) and molar refractive index (R_M) of Li-Ox, Na-Ox, K-Ox at different mass fractions of aqueous UA solution at 298.15K[a]	242

LIST OF FIGURES

CHAPTERS		FIGURES	PAGE NO.
CHAPTER-I	**Fig.I.a.**	Solvation of a salt.	20
	Fig.I.b.	Inclusion complex.	21
CHAPTER-II	**Fig.II.a.**	Pictorial representation of inclusion complex formation.	25
	Fig.II.b.	Pictorial diagram of (i) Inclusion complex (ii) Cucurbituril[6], (iii) 18-Crown-6 and (iv) P-tert-butylcalix-4-arene.	26
	Fig.II.c.	(a) Torus shape of cyclodextrin with dimensions and molecular structure of (b) α-cyclodextrin(c) β-cyclodextrin (d) γ-cyclodextrin.	28
	Fig.II.d.	Different stoichiometric ratios in host-guest inclusion complexes.	30
CHAPTER-IV	**Fig.1.**	Plot of molar conductance (Λ) and the square root ofconcentration ($\sqrt{c}$) of $LiAsF_6$ in w_1=0.00 (-♦-), w_1=0.25 (-■-), w_1=0.50 (-▲-), w_1=0.75 (-●-), and w_1=1.00 (-∗-) mass fraction of EG in MeOH at T=298.15K.	103
	Fig.2.	Plot of limiting molar conductance (Λ_0) (-♦-) and Walden Product ($\Lambda_0\eta$) (-▲-) of $LiAsF_6$ in different mass fraction (w_1) of EG in MeOH at T=298.15K.	103
	Fig.3.	Plot of limiting ionic conductance ($\lambda_o^{\pm}$) of Li^+ (-♦-), AsF_6^- (-■-) and ionic Walden Product ($\lambda_o^{\pm}\eta$) of Li^+ (-▲-), AsF_6^- (-●-) in different mass fraction (w_1) of EG in MeOH at T=298.15K.	105
	Fig. 4.	Plot of association constant (K_a) (-♦-), and Gibb's energy change (ΔG^0) (-▲-) of $LiAsF_6$ in different mass fraction (w_1) of EG in MeOH at 298.15K.	105
	Fig.5.	Plot of limiting apparent molar volume (ϕ_V^0) (-♦-) and viscosity B-coefficient (-▲-) vs. mass fraction (w_1) of EG in MeOH for $LiAsF_6$ at T=298.15K.	109
	Fig.6.	The plausible mechanism of interaction between $LiAsF_6$ and Solvents (ethylene glycol and methanol).	110
	Fig.7.	The plausible diagrams of molecular interaction and solvation of the ions in solvent mixtures.	112

CHAPTERS	FIGURES		PAGE NO.
CHAPTER-V	**Fig.1**	Representation of Molecular structure of of [BMIm]Cl and [HMIm]Cl and an overall view of cyclodextrin.	115
	Fig.2.	1H NMR spectra of (a) pure β –CD(top), [BMIm]Cl (middle) and β-CD&[BMIm]Cl complex (bottom) (b) pure β –CD(top), [HMIm]Cl (middle) and β –CD&[HMIm]Cl complex (bottom).	119
	Fig.3.	Plausible stoichiometric inclusion ratio of the host and guest molecule. (Host:Guest).	120
	Fig.4.	Linear relationships of logKa vs. 1/T for the interaction of [HMIm]Cl and α-CD (♦), [HMIm]Cl and β-CD (■), [BMIm]Cl and α-CD (▲), [BMIm]Cl and β-CD(•).	123
	Fig.5.	FT-IR Spectra of (a) pure β –CD, [BMIm]Cl, and [BMIm]Cl + β –CD inclusion complex (b)pure β –CD, [HMIm]Cl, [HMIm]Cl + β –CD inclusion complex.	125
	Fig.6	(a)Limiting apparent molar volume (φv0) and Sv* (b) viscosity A and B-coefficient of [BMIm]Cl and [HMIm]Cl in different mass fractions (w1) of aqueous (■) α-CD and (■) β-CD solution at 298.15K.	128
	Fig.7.	Schematic representation of inclusion of (a) [BMIm]Cl and [HMIm]Cl with α-CD and (b) [BMIm]Cl and [HMIm]Cl β-CD	131
	FigS1.	1H NMR spectra of (a) pure α–CD(top), [BMIm]Cl (middle) and α–CD&[BMIm]Cl complex (bottom) (b) pure α –CD(top), [HMIm]Cl (middle) and α–CD&[HMIm]Cl complex (bottom).	140
	Fig.S2.	Surface tension of (a) [BMIm]Cl with α –CD (b) [BMIm]Cl with β–CD (c) [HMIm]Cl with α –CD (D) [HMIm]Cl with β - CD at 298.15K.	141
	Fig.S3.	Plot of conductivity against square root of molar concentration (c) for (a) [BMIm]Cl + α-CD, (b) [BMIm]Cl+β-CD (c)[HMIm]Cl+α-CD and (d) [HMIm]Cl+β-CD complexes at 293.15 K (♦), 303.15K (■) and 313.15K (▲).	141
	FigS4.	Schematic representation of fate of the ionic liquid in aqueous medium.	142
	FigS5.	FT-IR Spectra of (a) pure β –CD, [BMIm]Cl, and [BMIm]Cl + β –CD inclusion complex (b)pure β –CD, [HMIm]Cl, [HMIm]Cl + β –CD inclusion complex.	143

CHAPTERS	FIGURES		PAGE NO.
	Fig.S6.	ESI Mass spectra of (a) [BMIm]Cl-α-CD (b) [BMIm]Cl-β-CD (c) [HMIm]Cl-α-CD (d) [HMIm]Cl-β-CD inclusion complexes.	145
CHAPTER-VI	**Fig.1.**	Job's Plot for (a) [BMPy]PF6 with α-CD and (b) [BMPy]PF6 with β-CD.	149
	Fig.2.	1H NMR spectra of pure α-CD, β–CD, [BMPy]PF6 and [BMPy]PF6+α-CD and [BMPy]PF6+β–CD. (In D2O, 400 MHz)	151
	Fig.3.	**2D** ROESY spectra of the solid ICs of [BMPy]PF6-α-CD AND [BMPy]PF6-β-CD in D2O. (Cross correlations are indicated by red circles)	153
	Fig.4.	Surface tension of [BMPy]PF6 with α –CD, [BMPy]PF6 with β–CD at 298.15 K.	154
	Fig.5.	Plot of log K_a^c vs 1/T for the interaction of α-CD with [BMPy]PF6 (♦) and β-CD with [BMPy]PF6(■).	157
	Fig.6.	Plot of log K_a vs 1/T for the interaction of α-CD with [BMPy]PF6 (♦) and β-CD with [BMPy]PF6(■).	158
	Fig.7:	FT-IR spectra of free α-CD, β-CD, [BMPy]PF6 and their 1:1incluusion complexes ([BMPy]PF6+α-CD) and ([BMPy]PF6+β-CD) at 298.15K.	161
	Fig.8.	Schematic representation of the inclusion of [BMPy]PF6 with Cyclodextrin in mixed solvent media.	162
	Fig.S1.	2D ROESY spectra of the solid ICs of [BMPy]PF6-α-CD in D2O. (Cross correlations are indicated byred circles)	172
	Fig.S2.	Plot of conductivity with increasing concentration of α-CD and β-CD with [BMPy]PF6+α-CD and [BMPy]PF6+β-CD at (a) 293.15 K, (b) 303.15K and (c) 313.15K.	173
	Fig.S3.	Benesi-Hildebrand double reciprocal plots for the effect of β-CD on theabsorbance of [BMPy]PF6 +β-CD and [BMPy]PF6+α-CD (252 nm) at different temperatures(a) and (b) shows at 293015K, (c) and (d) at 303.15K and (e) and (f) at 313.15K.	174
	Fig.S4.	ESI mass spectra of [BMPy]PF6-β-CD and [BMPy]PF6-α-CD inclusion complex.	175
CHAPTER-VII	**Fig.1.**	Molecular structures of (a) tetrabutylphosphonium methanesulfonate and (b) cyclodextrin molecule with interior and exterior protons (n = 6, 7 for α-CD and β-CD respectively).	178

CHAPTERS	FIGURES		PAGE NO.
	Fig.2.	1H-NMR spectra of (a) pure TBPMS (b) pure α–CD and (c) α–CD & TBPMS complex.	181
	Fig.3.	2D ROESY spectra of solid inclusion complex of 1:1 molar ratio of TBPMS+ α -CD complex and TBPMS + β-CD complex in D2O (correlation signals are marked by green borders).	183
	Fig.4.	Variation of surface tension of aqueous TBPMS with increasing concentration of aqueous α-CD and β-CD solution respectively at 298.15 K.	184
	Fig.5.	FTIR spectra of free α-CD, TBPMS and their 1:1 inclusion complex (TBPMS:α-CD).	187
	Fig.6.	Linear relationships of logKf vs. 1/T for the interaction of TBPMS+α-CD (▲), TBPMS+β-CD (♦).	189
	Fig.7.	Schematic representation of plausible interactions taking place in the inclusion complex.	191
	Fig.S1.	1H NMR spectra of (a) pure TBPMS (b) pure β –CD and (c) β–CD & TBPMS complex.	197
	Fig.S2.	Plot of conductivity against square root of molar concentration (√C) for TBPMS +α-CD(▲), TBPMS+β-CD(♦) at 293.15 K, 303.15K and 313.15K.	198
	Fig.S3.	FT-IR spectra of free β-CD, TBPMS and their 1:1 inclusion complex (TBPMS:β-CD).	199
	Fig.S4.	ESI mass spectra of TBPMS-β-CD inclusion complex and TBPMS-β-CD inclusion complex.	200
CHAPTER-VIII	**Fig.1.**	Molecular structure of (a) one unit of CD, (b) PTX and (c) 2-PAM.	203
	Fig.2.	Surface tension of PTX with α –CD (top left), PTX with β–CD (top right) 2-PAM with α –CD (bottom left) 2-PAM with β – CD (bottom right) at 298.15 K.	206
	Fig.3.	Job plot of (a) PTX-α-CD at λmax=272 nm, (b) PTX-β-CD at λmax=272 nm, (c) 2-PAM-α-CD at λmax=292 nm, (d) 2-PAM-β-CD at λmax=292 nm. ΔA= (AIC-ACD), R= [Drug]/([Drug]+[CD]) (Total concentration of the Soln. =100 μM).	208
	Fig.4.	Schematic diagrams of inclusion (a) 2-PAM with α-CD and β-CD, (b) PTX with 2-PAM with α-CD and β-CD.	210

CHAPTERS	FIGURES		PAGE NO.
	Fig.5	Schematic diagrams of inclusion through the narrower and wider rim of the cyclodextrin molecule.	211
	Fig.6	1H NMR spectra of (a) pure β –CD, PTX and β-CD + PTX (b) 2-PAM, pure β –CD and β-CD + PTX. (Green and blue colored protons are of PTX and 2-PAM respectively while red colored protons are of CD).	212
	Fig.7	FT-IR Spectra of (a) pure β –CD, PTX, and β-CD + PTX and (b) 2-PAM, pure β –CD, β-CD + PTX.	214
	Fig.8	(a) Limiting apparent molar volume (φv0) and (b) viscosity B-coefficient of PTX and 2-PAM in different mole fractions (w1) of aqueous (■) α-CD and (■)β-CD solution at 298.15K.	217
	Fig.S1.	Absorbance vs. wave length curve of α-CD and β-CD with PTX and 2-PAM.	226
	Fig.S2.	1H NMR spectra of (a) pureα–CD, PTX andα-CD + PTX; (b) 2-PAM, pure α –CD and α-CD + PTX; (Green and blue colored protons are of PTX and 2-PAM respectively while red colored protons are of CD).	227
	Fig.S3.	FT-IR Spectra of (a) pure β –CD, PTX, and β-CD + PTX and (b) 2-PAM, pure β –CD, β-CD + PTX.	228
CHAPTER-IX	**Fig.1.**	Molecular Structure of Uric acid.	231
	Fig.2.	Schematic diagram of (a) Potassium oxalate, (b) Sodium oxalate, (c) Lithium oxalate.	231
	Fig.3.	Plot of viscosity B-coefficient of Li-Ox (♦), Na-Ox (■), K-Ox (▼) at 293.15K (A), 303.15K (B), 313.15K (C) in aqueous Uric acid (UA) solution.	241
	Fig.4.	Plausible solvation consequences of Li+, Na+ and K+ (hydrated) in aqueous UA medium.	243