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The survey of literature on ethyl formate and ethyl chloroformate reveals that much attention has not been paid to study the solvent behaviour of these two compounds so far. The present work is being undertaken with a view to studying the potentialities of these two esters as nonaqueous solvents and to study the effect of replacement of hydrogen atom of formyl group in ethyl formate, with chlorine atom on the solvent behaviour of the ester. It is hoped that such a study will throw light on the chemistry of these two compounds. The present investigations may be described under the following headings:

a) Determination of solubility of various substances in the esters.
b) Formation of solvates in these esters.
c) Conductance measurements of solutions of Lewis acids, protonic acids and bases in the two esters.
d) Infrared spectral study of the structure of solvates.
e) Solvolytic reactions in these solvents.
f) Conductometric and potentiometric study of acid-base neutralisation reactions.

On the basis of evidence obtained from the above study probable modes of ionisation may be assigned to ethyl formate and ethyl chloroformate.